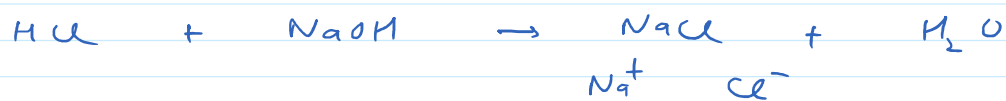


Family of salts

Salt is an ionic compound that can be formed by the neutralization reaction of an acid and a base. Salts are composed of related number of cations (positively charged ions) and anions (negatively charged ions) so that the product is electrically neutral (without a net charge).

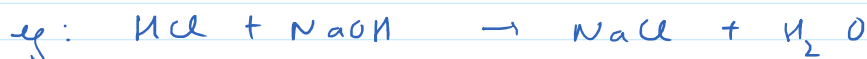
eg.



pH of salts

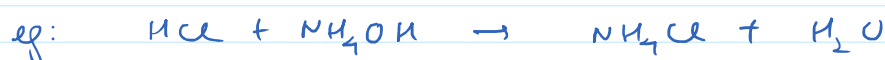
1. Neutral salts: (pH = 7)

Salts of strong acid and strong base are neutral with pH = 7.



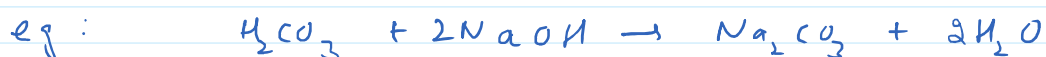
2. Acidic salt (pH < 7)

Salts of strong acid and weak base are acidic with pH < 7.



3. Basic salt (pH > 7)

Salt of strong base and weak acid are basic salts with pH > 7.



Salt	pH	Acid used	Base used
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Salt	pH	Acid used	Base used
NaCl (Na^+ Cl^-)	7	HCl	NaOH
KNO_3 (K^+ NO_3^-)	7	HNO_3	KOH
AlCl_3 (Al^{3+} Cl^-)	< 7	HCl	$\text{Al}(\text{OH})_3$
$\underline{\text{ZnSO}_4}$ (Zn^{+2} SO_4^{2-})	< 7	H_2SO_4	$\text{Zn}(\text{OH})_2$
CuSO_4 (Cu^{2+} SO_4^{2-})	< 7	H_2SO_4	$\text{Cu}(\text{OH})_2$
CH_3COONa (Na^+ CH_3COO^-)	> 7	CH_3COOH	NaOH
Na_2CO_3 (Na^+ CO_3^{2-})	> 7	H_2CO_3	NaOH
NaHCO_3 (Na^+ HCO_3^-)	> 7	H_2CO_3	NaOH