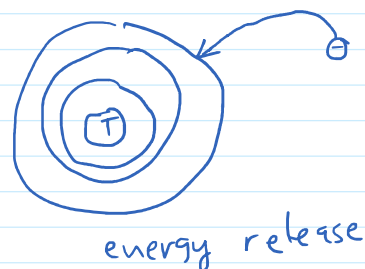
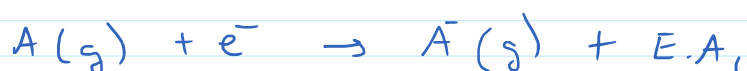


5. Electron affinity

The amount of energy released when an e^- is added to an isolated neutral gaseous atom in its ground state to produce an anion.

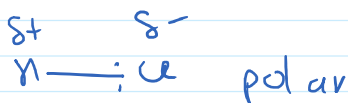
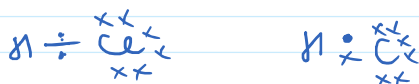


- Across a period from left to right electron affinity increases because incoming electron is pulled more strongly, hence more energy is released.
- Down the group electron affinity decreases because incoming electron is added in a shell farther away from the nucleus, hence protonic pull is lesser and lesser energy is released.

6. Electronegativity.

It is relative tendency of an atom of an element present in a molecule to attract the shared pair of e^- towards itself.

HCl



Across a period from left to right Electronega-

-tivity increase

Down the group electronegativity decreases