

Properties of ionic compounds

i) Physical nature

Ionic compounds are hard solids because of strong force of attraction between positive and negative ions. These compounds are brittle.

ii) Melting and boiling point:

These have high melting and boiling points. This is because a considerable amount of energy is required to break the strong inter-ionic attractions.

iii) Solubility

Electrovalent compounds are generally soluble in water and insoluble in solvents such as kerosene, petrol etc.

iv) Conduction of electricity

The conduction of electricity through a solution involves the movement of charged particles.

In molten state and in aqueous state ionic compounds conduct electricity due to presence of free ions.

In solid state ionic compounds do not conduct electricity because ions are not free as they are held strongly with electrostatic forces of attraction.