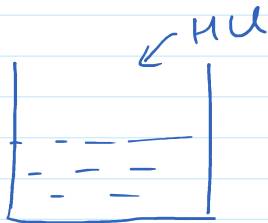


At 25°C $[H^+][OH^-] = \text{constant} = 10^{-14}$



water

$[H^+] = 1M = 10^0 M$

$[H^+] = 10^{-pH}$

$[H^+]$	10^0	10^{-1}	10^{-2}	10^{-3}	10^{-4}	10^{-5}	10^{-6}	10^{-7}
$[OH^-]$	10^{-14}	10^{-13}	10^{-12}	10^{-11}	10^{-10}	10^{-9}	10^{-8}	10^{-7}
pH	0	1	2	3	4	5	6	7

$\xrightarrow{\text{[H}^+] \text{ decrease}}$ $\xleftarrow{\text{[OH}^-] \text{ increase}}$
 acidic neutral

$[H^+]$	10^{-8}	10^{-9}	10^{-10}	10^{-11}	10^{-12}	10^{-13}	10^{-14}
$[OH^-]$	10^{-6}	10^{-5}	10^{-4}	10^{-3}	10^{-2}	10^{-1}	10^0
pH	8	9	10	11	12	13	14

$[OH^-] > [H^+]$

basic