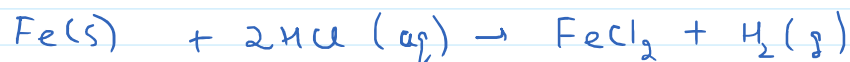
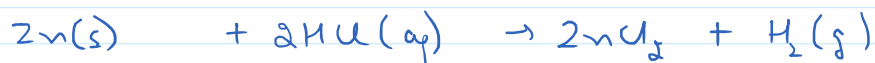
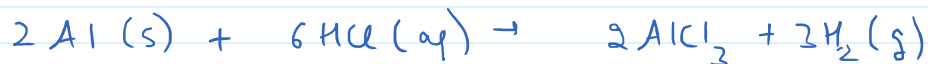
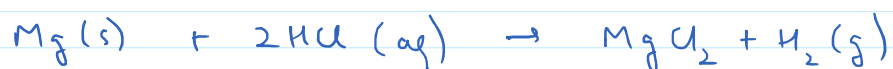


Reaction of metals with acids

i) Metal + dilute acid \rightarrow salt + hydrogen



ii) Metals below hydrogen in reactivity series do not react with dilute HCl.

eg: Cu, Hg, Ag, Au.

iii) Hydrogen gas is not evolved when a metal reacts with nitric acid because HNO_3 is a strong oxidising agent. It oxidises H_2 produced to water and itself gets reduced to any of the nitrogen oxides (N_2O , NO , NO_2)

iv) Mg, Mn react with very dilute HNO_3 to evolve H_2 gas

$\begin{matrix} \text{H} & \text{H} & \text{H} \\ \text{Hydrogen} & \text{Mg} & \text{Mn} \end{matrix}$