

Terms related to electrolytic conduction

i) Resistance (R)

Hindrance to the flow of ions

SI unit: ohm, Ω

ii) Conductance (C)

Easiness with which ions can flow, it is reciprocal of resistance.

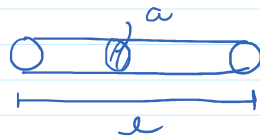
$$C = \frac{1}{R}, \text{ unit} = \text{ohm}^{-1}, \text{ mho}, \text{ Seimen (S)}$$

iii) Resistivity (ρ) or Specific resistance

$$R = \rho \frac{l}{a}$$

$l \rightarrow$ length of section through which current flows

$a \rightarrow$ area of cross section



$$\rho = R \frac{a}{l} \quad \text{units: ohm-cm}$$

$$a = 1 \text{ unit}, l = 1 \text{ unit}$$

$$\rho = R$$

$$a = 1 \text{ cm}^2, l = 1 \text{ cm}, \text{ volume} = 1 \text{ cm}^3 = 1 \text{ ml}$$

Resistivity is defined as resistance offered to all the ions present in 1 ml solution.

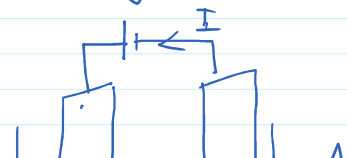
iv) Conductivity (k) or specific conductance

It is conductance of 1 ml solution

$$k = \frac{1}{\rho}$$

$$\text{units: } \frac{1}{\text{ohm cm}} = \text{mho cm}^{-1} = \text{Scm}^{-1}$$

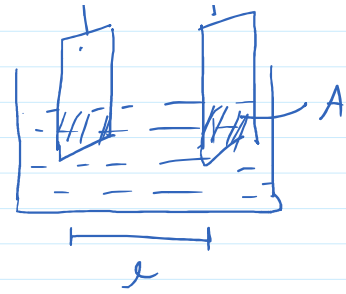
Cell constant



Cell constant

$$\text{Cell constant } (G) = \frac{l}{a}$$

unit : cm^{-1}



$$R = \rho \frac{l}{a}$$
$$= \rho G$$

$$G = \frac{R}{\rho}$$

$$G = R \times K$$

cell constant = resistance \times conductivity

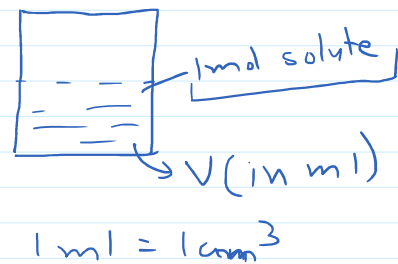
ii) Molar conductivity

Conductance of all the ions produced by 1 mol of electrolyte is known as molar conductivity.

$$\Lambda_m = K \times V$$

units : $\text{Scm}^{-1} \frac{\text{ml}}{\text{mol}}$

$$= \text{Scm}^{-1} \frac{\text{cm}^3}{\text{mol}}$$



unit = $\text{Scm}^2 \text{mol}^{-1}$

Let a solution has molarity M

$$\begin{array}{ccccccc} M \text{ moles electrolyte} & \text{is present in} & = & \frac{1000 \text{ ml}}{M} & \text{solution} \\ 1 \text{ mole} & & & & = & \frac{1000 \text{ ml}}{M} & \text{solution} \end{array}$$

$$V = \frac{1000 \text{ ml}}{M}$$

$$\Lambda_m = K \times V$$

$$\Lambda_m = \frac{K \times 1000}{M} \quad \text{units: } \text{Scm}^2 \text{mol}^{-1}$$

iii) Equivalent conductivity (Λ_{eq})

Conductance of solution containing 1 equivalent of electrolyte

$$\Lambda_{eq} = K \times V$$

v = volume of solution in ml containing 1 equivalent of electrolyte.

$$\Lambda_{eq} = \frac{K \times 1000}{N}$$

units $\text{Scm}^2 \text{eq}^{-1}$

$$N = n \text{ factor} \times M$$

$$\Lambda_{eq} = \frac{K \times 1000}{n \text{ factor} \times M}$$

$$\Lambda_{eq} = \frac{\Lambda_m}{n \text{ factor}}$$