

Mendeleev's Periodic Table

- Dmitri Ivanovich Mendeleev is known as the father of Modern Periodic Table.
- When he started his work, 63 elements were known.
- He examined the relationship between atomic masses of elements and their physical and chemical properties.
- Among the chemical properties, he concentrated on the compounds formed by elements with oxygen and hydrogen as H_2 and O_2 are very reactive and formed compounds with most elements.
- He took 63 cards and on each card he wrote down the properties of one element.
- He sorted out the elements with similar properties and pinned the cards together on a wall.
- He observed that cards were arranged in increasing order of atomic masses, also there was periodic recurrence of elements with similar chemical and physical properties.
- On the basis of this he formulated a law "The properties of elements are the periodic function of their atomic masses."
- He formed 8 groups (vertical lines)
6 periods (horizontal lines)
- After the discovery of noble gases and radioactive elements, another version of Mendeleev's periodic table was published in 1905.

Group zero was added for noble gases

7th period was added for radioactive elements.